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~~to Mole, Grams to~~

~~Grams, Mole Ratio~~

~~Practice Problems~~

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~~What is~~

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Introduction to
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How to Calculate
Limiting Reactant
and Moles of
Product

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Table

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to Chemistry, Basic
Concepts - Periodic

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of Chemical

Reactions ALL OF
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the answer on the
line to the left.

Show all your work
in the space
provided. 1. _____

The actual yield of
a reaction is 22 g

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and the theoretical
yield is 25 g.

Calculate the
percentage yield. 2.

6.0 mol of N_2 are
mixed with 12.0

mol of H
CHAPTER
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Reaction

Stoichiometry. Mole

Ratio. Limiting

Reactant.

Calculations

involving the mass

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relationships of
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Calculations
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between reactant...

A conversion factor
that relates the
amounts in moles of
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controls the
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ANSWER Answer

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The coefficients in a chemical equation represent the (a)

masses in grams of all reactants and products. (b)

relative number of moles of reactants and products. (c)

number of atoms of each element in each compound in a reaction. (d)

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number of valence
electrons involved
in a reaction ...

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Stoichiometry Ideal

Stoichiometric

Calculations

Limiting Reactant

and Percent Yield 2

Definitions.

Composition

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Stoichiometry deals with the mass relationships of elements in compounds.

Reaction

stoichiometry involves the mass relationships between reactants and products in a chemical reaction.

3. Definitions. A mole ratio is ...

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stoichiometry uses
molar relationships

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to determine the amounts of unknown reactants or products from the amounts of known reactants or products.

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involves the mass relationships between reactants and products in a chemical reaction.

Reaction

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