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Worksheet 2

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~~Dilutions~~

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Formula \u0026amp;

Equations

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Dilution

Problems

Solution

Stoichiometry

Grams, Moles,

Liters Volume

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Chemistry

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Practice

Problems

Molarity

Practice

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Chemistry: How
to Calculate and
Perform Molarity
Dilutions *Mole*
Conversions Made

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Easy: How to
Convert Between
Grams and Moles
~~Molarity and~~
~~Dilution~~
~~Worksheet~~
~~Molarity,~~
~~Solutions,~~
~~Concentrations~~
~~and Dilutions~~

Molarity
Practice
Problems (Part
2) dilution

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calculation 2

*Solubility Rules
and How to Use a
Solubility Table*

Step by Step

Stoichiometry

Practice

Problems | How
to Pass

Chemistry

Dilution Series

\u0026amp; Serial

Dilution

Molarity Made

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**Easy: How to
Calculate
Molarity and
Make Solutions**

Oxidation and

Reduction

(Redox)

Reactions Step-

by-Step Example

Molarity

Problems and

Examples

Percentage

Concentration

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Calculations

~~Dilution~~

~~Problems~~

Stoichiometry:

Converting Grams

to Grams

Dilution

Explained How to

Do Solution

Stoichiometry

Using Molarity

as a Conversion

Factor | How to

Pass Chemistry

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Solution
Stoichiometry
and Dilution

Answers

Preparing
Solutions - Part
3: Dilutions
from stock
solutions

**Mass
Percent \u0026
Volume Percent -
Solution**

Composition

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Chemistry

The Dilution

Problems

~~Molarity:~~

~~Dilution~~

~~Practice Problem~~

~~2 Molarity and~~

~~Dilution L-2 ||~~

MOLARITY ,

DILUTION \u0026

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Osmosis and

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Worksheet 2

(Updated)

Molarity

Worksheet 2 The

Dilution

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Problems

Worksheet $M=nV$

n = # moles V

must be in

liters (change

if necessary) 1.

What is the

molarity of a

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Worksheet 2

0.30 liter solution containing 0.50 moles of NaCl?

2. Calculate the molarity of 0.289 moles of FeCl_3 dissolved in 120 ml of solution? 3. If a 0.075 liter solution c...

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Dilutions 2

Worksheet -

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Dilution

Problems

Worksheet 1. How do you prepare a 250.-ml of a 2.35 M HF dilution from a 15.0 M stock solution? 2. If 455-ml of 6.0 M HNO_3 is used to

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Worksheet 2
make a 2.5 L
dilution, what
is the molarity
of the dilution?
3. If 65.5 ml of
HCl stock
solution is used
to make 450.-ml
of a 0.675 M HCl
dilution, what
is

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Problems

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Worksheet – Mrs

Getson's Blog

Dilutions

Worksheet –

Solutions 1) If

45 mL of water
are added to 250
mL of a 0.75 M K_2SO_4 solution,
what will the
molarity of the
diluted solution
be? $(0.75 \text{ M})(250 \text{ mL}) = M_2(295$

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mL) $M_2 = (0.75$

M) (250 mL) =

0.64 M (295 mL)

2) If water is

added to 175 mL

of a 0.45 M KOH

solution until

the volume is

250 mL, what

Dilutions

Worksheet W 329

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Solutions

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Dilution

Problems Answer

Key Program 1)

If 45 mL of

water are added

to 250 mL of a

0.75 M K_2SO_4

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Worksheet 2
solution, what
will the
molarity of the
diluted solution
be? 2) If water
is added to 175
mL of a 0.45 M

Solutions

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Dilution

Problems ...

Dilutions

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Worksheet 2

Solutions 1) If I add 25 mL of water to 125 mL of a 0.15 M NaOH solution, what will the molarity of the diluted solution be? $M_1V_1 = M_2V_2$

$(0.15 \text{ M})(125 \text{ mL}) = x(150 \text{ mL})$ $x = 0.125 \text{ M}$ 2) If I add water to 100

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Worksheet 2
mL of a 0.15 M
NaOH solution
until the final
volume is 150
mL, what will
the molarity of
the diluted
solution be?
M1V1 ...

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Problems

Worksheet $M=nV$

$n = \# \text{ moles}$ V

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liters (change if necessary) 1.

What is the molarity of a

0.30 liter solution

containing 0.50 moles of NaCl?

2. Calculate the molarity of

0.289 moles of FeCl₃ dissolved in 120 ml of

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Solutions 1) If
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are added to 250
mL of a 0.75 M K_2SO_4 solution,
what will the
molarity of the
diluted solution
be? $(0.75 \text{ M})(250 \text{ mL}) = M_2(295 \text{ mL})$
 $M_2 = (0.75 \text{ M})(250 \text{ mL}) =$
 $0.64 \text{ M}(295 \text{ mL})$
2) If water is

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added to 175 mL
of a 0.45 M KOH
solution until
the volume is
250 mL, what

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Problems Answer

Key dissolved to

make 0.10 L of

solution. 2) 1.0

grams of

potassium

fluoride is

dissolved to

make 0.10 L of

solution.

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Dilution Worksheet 2

Problems Answers

where the subscripts "1" and "2" refer to the solution before and after the dilution, respectively.

Since the dilution process does not change the amount of solute in the

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Worksheet 2
solution, $n_1 = n_2$

2. Thus, these two equations may be set equal to one another:

$$[M_1L_1 = M_2L_2]$$

This relation is commonly referred to as the dilution equation.

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